## Standard Atomic Notation Worksheet

## Recall:

- mass number - atomic number = number of neutrons
- atomic number $=$ number of protons
- standard atomic notation:

$$
\begin{array}{cc}
\text { Atomic Mass } \rightarrow & 12 \\
\text { Atomic Number } & 6
\end{array}
$$

Given the following information, determine the standard atomic notation for the element:

1) Element: Calcium, Mass number: 40, Atomic Number: 20
\# of Neutrons =
\# of Protons =
\# of Electrons =

2) Element: Chlorine, Mass number: 36, Atomic Number: 17 \# of Neutrons =
\# of Protons =
\# of Electrons =

3) Element: ???, Mass number: 65, Atomic Number: 30
\# of Neutrons =
\# of Protons =
\# of Electrons =

4) Element: Gold, Mass number: 197, Atomic number: ??
\# of Neutrons =
\# of Protons =
\# of Electrons =

5) Pick your own element: $\qquad$ Mass number: $\qquad$ Atomic number: $\qquad$ \# of Neutrons =
\# of Protons = \# of Electrons =

